C.U.SHAH UNIVERSITY
Winter Examination-2018

Subject Name: Pharmaceutical Analysis-I

Subject Code: 4PS02PHA1   **Branch:** B. Pharm

Semester: 2   **Date: 25/10/2018**   **Time: 02:30 To 05:30**   **Marks:** 70

Instructions:
  (1) Use of Programmable calculator & any other electronic instrument is prohibited.
  (2) Instructions written on main answer book are strictly to be obeyed.
  (3) Draw neat diagrams and figures (if necessary) at right places.
  (4) Assume suitable data if needed.

Q-1   Attempt the following questions: **(14)**

a)	Define Accuracy.	(1)
b)	Define Normality.	(1)
c)	Give one example of Self indicator.	(1)
d)	Define pH and write its range.	(1)
e)	Define Supersaturation.	(1)
f)	Write the color change of phenolphthalein in different medium.	(1)
g)	Define Common ion effect.	(1)
h)	Define Molarity.	(1)
i)	Give Examples of Buffer solution.	(1)
j)	Define Iodometry.	(1)
k)	Write example of Complexometric titration.	(1)
l)	Define Primary standard.	(1)
m)	Define Lewis acid.	(1)
n)	Define Error.	(1)

Attempt any four questions from Q-2 to Q-8

Q-2   **(14)**
a   Write the theory of Acid Base indicator. Describe.   (7)
b   Describe steps involved in Gravimetric Analysis.   (7)

Q-3   **(14)**
a   Explain Fajan's method of analysis in detail.   (7)
b   Define Solubility product and explain effect of Acid, Temperature and Solvent on   (7)
         solubility product.

Q-4 **(14)**

   a   Describe Law of mass action. (7)

   b   Write detail application of Complexometric titration. (7)

Q-5 **(14)**

   a   Enumerate types of Error and methods of minimizing error. (7)

   b   Explain neutralization curve of Acid base titration. (7)

Q-6 **(14)**

   a   Define Complexometric titration and write detail note on Metal ion indicator. (7)

   b   Explain Hydrolysis of Salt. (7)

Q-7 **(14)**

   a   Explain Co-precipitation along with its types and differentiate between co-precipitates and post precipitate. (7)

   b   Define buffer with explain and explain Henderson-Hessselblach equation. (7)

Q-8 **(14)**

   a   Describe Acidimetry and Alkalimetry in nonaqueos titration. (7)

   b   Explain Cerrimetry in detail (7)